ગુજરાત રાજ્યના શિક્ષણવિભાગના પત્ર-ક્રમાંક
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CHEMISTRY
PRACTICAL RECORD BOOK
(JOURNAL)

Standard XI

PLEDGE

India is my country.
All Indians are my brothers and sisters.
I love my country and I am proud of its rich and varied heritage.
I shall always strive to be worthy of it.
I shall respect my parents, teachers and all my elders and treat everyone with courtesy.
I pledge my devotion to my country and its people.
My happiness lies in their well-being and prosperity.

રાજ્ય સરકારની વિનામૂલ્યે યોજના હેઠળનું પુસ્તક

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PREFACE

In accordance with the Government's Policy to implement uniformed curriculum at national level, the Government of Gujarat and Secondary and Higher Secondary Board of Education, Gujarat State have decided to implement the NCERT Textbooks at school level directly as per resolution no M.SH.B./1217/1036/CHH dated 25-10-17. With reference to that Gujarat State Board of School Textbooks is pleased to introduce this **Chemistry Practical Record Book (Journal)** prepared for **Standard XI** before the students.

Before publishing the Chemistry Practical Record Book (Journal) its manuscript has been fully reviewed by expert professors and teachers. According to their suggestions, we have made necessary changes in the manuscript before publishing the Chemistry Practical Record Book (Journal). The Board has taken special care to ensure that this Chemistry Practical Record Book (Journal) is made interesting, useful and free from errors. However, to improve the quality of the Chemistry Practical Record Book (Journal), we welcome suggestions, from people interested in education.

P. Bharathi (IAS)
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Date : 09-12-2019

Executive President
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FUNDAMENTAL DUTIES

It shall be the duty of every citizen of India :*

(a) To abide by the Constitution and respect its ideals and institutions, the National Flag and the National Anthem;

(b) To cherish and follow the noble ideals which inspired our national struggle for freedom;

(c) To uphold and protect the sovereignty, unity and integrity of India;

(d) To defend the country and render national service when called upon to do so;

(e) To promote harmony and the spirit of common brotherhood amongst all the people of India transcending religions, linguistic and regional or sectional diversities; to renounce practices derogatory to the dignity of women;

(f) To value and preserve the rich heritage of our composite culture;

(g) To protect and improve the natural environment including forests, lakes, rivers and wildlife, and to have compassion for living creatures;

(h) To develop the scientific temper, humanism and the spirit of inquiry and reform;

(i) To safeguard public property and to abjure violence;

(j) To strive towards excellence in all spheres of individual and collective activity so that the nation constantly rises to higher levels of endeavour and achievement.

(k) who is a parent or guardian to provide opportunities for education to his child or, as the case may be, ward between the age of six and fourteen years.

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* Constitution of India : Section 51-A

CERTIFICATE

Exam Seat No. : _____

CHEMISTRY PRACTICAL RECORD BOOK (JOURNAL)

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_____

has performed practical work of _____ chemistry practicals out of _____ practicals satisfactorily.

_____ _____

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Date : _____

Experiment 1

Aim : Preparation of _____ mL of 0.1M standard solution of oxalic acid.

[Reference: Experiment No. 2.1 of Laboratory Manual]

Apparatus Required :

_____

_____

Materials Required :

_____

Calculation :

Standard solution will be prepared according to the experiment procedure by using the required amount of compound in gram calculated with the help of following formula.

$$\text{Weight of compound} = \frac{\text{Molarity} \times \text{Molar mass} \times \text{Volume of solution (mL)}}{1000}$$

$$\therefore \text{ Weight of oxalic acid } = \frac{\qquad\qquad}{1000}$$

$$= \_\_\_\_\_ \text{ g}$$

Procedure :

_____

_____

_____

_____

_____

_____

- **Answer the following questions :**

1. What is the formula and the basicity of hydrated oxalic acid and anhydrous oxalic acid ?

2. What do you mean by a molar solution ?

3. Why are the standard solutions always prepared in a volumetric flask ?

4. How will you prepare 250 mL of 0.05 M oxalic acid solution ?

5. Can solid NaOH be used to prepare its standard solution ?

6. What type of substance can be used for preparing standard solution ?

7. What is meant by "weighing by transfer"? When is this used ?

Sign of subject teacher/Lab teacher

Date : _____

Experiment 2

Aim : To determine the concentration (strength) of given dilute sodium hydroxide solution by titrating it against a 0.1 M oxalic acid solution.

[Reference: Experiment No. 6.1 of Laboratory Manual]

Apparatus Required :

_____

_____

Materials Required :

_____

_____

- Burette : _____
- Conical flask : _____
- Indicator : _____
- Colour change : _____
- Chemical Equation :

Procedure :

_____

_____

_____

_____

_____

Observation :

Burette Reading	I	II	III	IV	Concordant Reading (mL)
Final Reading (mL)					
Initial Reading (mL)					
Difference (mL)					

Calculation :

(1) Molar mass of NaOH :

(2) Determination of molarity of NaOH solution :

Sodium hydroxide Oxalic acid

$a_1M_1V_1 = a_2M_2V_2$

where a_1 = Acidity of _____ = _____

M_1 = Molarity of _____ = _____

V_1 = Volume of _____ = _____

a_2 = Basicity of _____ = _____

M_2 = Molarity of _____ = _____

V_2 = Volume of _____ = _____

$\therefore$ Molarity of NaOH solution = _____ M

(3) Concentration (strength) of NaOH solution = Molarity $\times$ Molar mass

= _____

= _____ g/L

Result :

Concentration of NaOH solution is _____ g/L

- **Answer the following questions :**

1. Why are the burette and the pipette rinsed with the solution with which these are filled ?

2. What is an indicator? Which indicator is used in the titration of oxalic acid vs sodium hydroxide? Can the titration be performed by using some other indicator ?

3. Why should one read the lower meniscus in the case of colourless and transparent solutions and the upper meniscus for solutions of dark colour ?

4. Explain the term 'end point'.

5. What do you mean by 1.0 M solution ?

6. Why should the last drop of the solution not be blown out of a pipette ?

7. Explain the term basicity of an acid and acidity of a base ?

8. For titrating NaOH vs HCl, phenolphthalein and methyl orange, both are suitable indicators. Why ?

9. What is meant by the term, 'concordant readings' ?

10. Can one take oxalic acid solution in the burette and sodium hydroxide solution in the titration (conical) flask ? Point out the limitations of doing so if any.

Sign of subject teacher/Lab teacher

Date : _____

Experiment 3

Aim : To determine the concentration (strength) of given dilute potassium hydroxide solution by titrating it against a 0.1 M oxalic acid solution.

[Reference : Experiment No. 6.1 of Laboratory Manual]

Apparatus Required :

_____

_____

Materials Required :

_____

_____

- Burette : _____
- Conical flask : _____
- Indicator : _____
- Colour change : _____
- Chemical Equation :

Procedure :

_____

_____

_____

_____

Observation :

Burette Reading	I	II	III	IV	Concordant Reading (mL)
Final Reading (mL)					
Initial Reading (mL)					
Difference (mL)					

Calculation :

(1) Molar mass of KOH :

(2) Determination of molarity of KOH solution :

Potassium hydroxide Oxalic acid

$a_1M_1V_1 = a_2M_2V_2$

where a_1 = Acidity of _____ = _____

M_1 = Molarity of _____ = _____

V_1 = Volume of _____ = _____

a_2 = Basicity of _____ = _____

M_2 = Molarity of _____ = _____

V_2 = Volume of _____ = _____

$\therefore$ Molarity of KOH solution = _____ M

(3) Concentration (strength) of KOH solution = Molarity $\times$ Molar mass

= _____

= _____ g/L

Result :

Concentration of KOH solution is _____ g/L

Sign of subject teacher/Lab teacher

Date : _____

Experiment 4

Aim : Preparation of _____ mL of 0.1M standard solution of sodium carbonate.

[Reference: Experiment No. 6.2 of Laboratory Manual]

Apparatus Required :

_____

_____

Materials Required :

_____

Calculation :

Standard solution will be prepared according to the experiment procedure by using the required amount of compound in gram calculated with the help of following formula.

$$\text{Weight of compound} = \frac{\text{Molarity} \times \text{Molar mass} \times \text{Volume of solution (mL)}}{1000}$$

$$\therefore \text{ Weight of sodium carbonate} = \frac{\qquad\qquad\qquad}{1000}$$

$$= \_\_\_\_\_ \text{ g}$$

Procedure :

_____

_____

_____

_____

_____

_____

Sign of subject teacher/Lab teacher

Date : _____

Experiment 5

Aim : To determine the concentration (strength) of given dilute hydrochloric acid solution by titrating it against a solution of 0.1 M sodium carbonate.

[Reference: Experiment No. 6.3 of Laboratory Manual]

Apparatus Required :

_____

_____

Materials Required :

_____

_____

- Burette : _____
- Conical flask : _____
- Indicator : _____
- Colour change : _____
- Chemical Equation :

Procedure :

_____

_____

_____

_____

Observation :

Burette Reading	I	II	III	IV	Concordant Reading (mL)
Final Reading (mL)					
Initial Reading (mL)					
Difference (mL)					

Calculation :

(1) Molar mass of HCl :

(2) Determination of molarity of HCl solution :

Hydrochloric acid Sodium carbonate

$a_1M_1V_1 = a_2M_2V_2$

where a_1 = Basicity of _____ = _____

M_1 = Molarity of _____ = _____

V_1 = Volume of _____ = _____

a_2 = Acidity of _____ = _____

M_2 = Molarity of _____ = _____

V_2 = Volume of _____ = _____

$\therefore$ Molarity of HCl solution = _____ M

(3) Concentration (strength) of HCl solution = Molarity $\times$ Molar mass

= _____

= _____ g/L

Result :

Concentration of HCl solution is _____ g/L

- **Answer the following questions :**

1. Which indicator is used in the titration of sodium carbonate against hydrochloric acid and what is the colour change at the end point ?

2. How will you prepare 250 mL of 0.05 M solution of sodium carbonate ?

3. Though sodium carbonate is a salt yet, its aqueous solution is weakly alkaline in nature. Explain why ?

4. How can you determine the acidity of sodium carbonate solution ?

5. Why is methyl orange not an Arrhenius base ?

6. How can you titrate a solution of the mixture of Na_2CO_3 and $NaHCO_3$ against HCl?

7. What is the difference between an end point and an equivalence point ?

8. Can you directly prepare standard solution of HCl, HNO_3, and H_2SO_4 ?

Sign of subject teacher/Lab teacher

Date : _____

Experiment 6

Aim : To determine the concentration (strength) of given dilute nitric acid solution by titrating it against a solution of 0.1 M sodium carbonate.

[Reference: Experiment No. 6.3 of Laboratory Manual]

Apparatus Required :

_____

_____

Materials Required :

_____

_____

- Burette : _____
- Conical flask : _____
- Indicator : _____
- Colour change : _____
- Chemical Equation :

Procedure :

_____

_____

_____

_____

Observation :

Burette Reading	I	II	III	IV	Concordant Reading (mL)
Final Reading (mL)					
Initial Reading (mL)					
Difference (mL)					

Calculation :

(1) Molar mass of HNO_3 :

(2) Determination of molarity of HNO_3 solution :

Nitric acid Sodium carbonate

$a_1M_1V_1 = a_2M_2V_2$

where a_1 = Basicity of _____ = _____

M_1 = Molarity of _____ = _____

V_1 = Volume of _____ = _____

a_2 = Acidity of _____ = _____

M_2 = Molarity of _____ = _____

V_2 = Volume of _____ = _____

$\therefore$ Molarity of HNO_3 solution = _____ M

(3) Concentration (strength) of HNO_3 solution = Molarity $\times$ Molar mass

= _____

= _____ g/L

Result :

Concentration of HNO_3 solution is _____ g/L

Sign of subject teacher/Lab teacher

Date : _____

Experiment 7

Aim : To determine the pH of some fruit juices.

[Reference: Experiment No. 5.1 of Laboratory Manual]

Apparatus Required :

Materials Required :

Procedure :

Observation :

pH values of different fruit juices

Name of the juice	Colour with universal indicator	pH	Colour of pH paper	pH	Inference (Acidic/Basic/Neutral)
Lemon					Shows _____ nature
Orange					Shows _____ nature
Apple					Shows _____ nature
Pineapple					Shows _____ nature

Result :

Increasing order of pH values of juices is _____ , _____ , _____ , _____
(Write names of juices)

- **Answer the following questions :**

1. Out of four juices, which one is least acidic? Explain.

_____

_____

2. If we dilute each of the juices, what effect is likely to be observed on the pH values ?

_____

_____

3. On mixing any two juices, would the pH alter or remain the same? Verify your answer experimentally ?

_____

_____

4. How can you ascertain the pH of a soft drink ?

_____

_____

Sign of subject teacher/Lab teacher

Date : _____

Experiment 8

Aim : To observe the variation in pH of acid/base with dilution.

[Reference: Experiment No. 5.2 of Laboratory Manual]

Apparatus Required :

Materials Required :

Procedure :

Observation :

pH change on dilution

Boiling Tube	HCl					H_2SO_4					NaOH				
	pH Paper		Universal Indicator		$[H^+]$	pH Paper		Universal Indicator		$[H^+]$	pH Paper		Universal Indicator		$[OH^-]$
	Colour	pH	Colour	pH		Colour	pH	Colour	pH		Colour	pH	Colour	pH	
A					0.1M					0.05M					0.1M
B															
C															
D															

Results :

(1) Concentration of H^+ of HCl and H_2SO_4 solution contained in boiling tubes B, C, and D is _____ (increasing /decreasing) respectively.

(2) Concentration of OH^- of NaOH solution contained in boiling tubes B, C, and D is _____ (increasing /decreasing) respectively.

Conclusion :

(1) On dilution of aqueous solution of acid, the concentration of hydrogen ion _____ (increases/decreases) and the pH values _____ (increases/decreases).

(2) On dilution of aqueous solution of base, the concentration of hydroxyl ion _____ (increases/decreases) and the pH values _____ (increases/decreases).

- **Answer the following questions :**

1. What trend is observed in the variation of pH with dilution for acidic as well as for basic solutions ?

2. How do you explain the results of variation in pH with dilution ?

3. If any two acidic solutions (say A and C) are mixed, what would happen to the pH of the mixture? Verify your answer experimentally.

4. For each acidic solution, whether we use HCl or H_2SO_4, pH is same to a reasonably good extent, even though HCl is 0.1M, and H_2SO_4 is 0.05M. How do you explain this result ?

5. Will the pH of 0.1M acetic acid be the same as that of 0.1M hydrochloric acid ? Verify your result and explain it.

Sign of subject teacher/Lab teacher

Date : _____

Experiment 9

Aim : To study the variation in pH by common ion effect in case of weak acids and weak bases.

[Reference: Experiment No. 5.3 of Laboratory Manual]

Apparatus Required :

Materials Required :

Procedure :

Observation :

Comparison of pH of acid/base and its buffer solution

Sl. No. of test Tube	Composition of the system	Colour of solution changed after adding universal indicator	pH
1.	CH_3COOH in water		
2.	NH_4OH (NH_3 in water)		
3.	$CH_3COOH + CH_3COONa$		
4.	$NH_4OH + NH_4Cl$		

Result:

(1) pH of aqueous solution of acetic acid (test tube 1) is _____ . (Write value)

(2) pH of buffer solution of acetic acid and sodium acetate (test tube 3) is _____ (higher/lower) than pH of acetic acid (test tube 1).

(3) pH of aqueous solution of ammonia (test tube 2) is _____ . (Write value)

(4) pH of buffer solution of aqueous ammonia solution and ammonium chloride (test tube 4) is _____ (higher/lower) than pH of ammonia solution (test tube 2).

Conclusion:

Common ion effect _____ (increases/decreases) the ionization of acid/base.

- **Answer the following questions :**

1. The addition of sodium acetate to acetic acid increases the pH whereas, the addition of NH_4Cl to aqueous NH_3 solution (NH_4OH) decreases the pH of the system. How do you explain these observations ?

2. Suggest suitable replacement for CH_3COONa for system 3 and NH_4Cl for system 4.

3. Suggest other pairs of weak acid and its salt and weak base and its salt to carry out the present investigations.

4. In salt analysis/mixture analysis, point out the situations where the variation in pH is carried out by common ion effect.

5. How do buffer solutions resist change in the pH ? Explain with a suitable example.

Sign of subject teacher/Lab teacher

Date : _____

Experiment 10

Aim : To study the change in pH during the titration of a strong acid with a strong base by using universal indicator.

[Reference: Experiment No. 5.4 of Laboratory Manual]

Apparatus Required :

Materials Required :

Procedure :

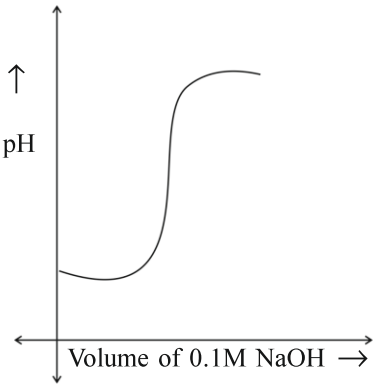
Observation :

pH change during the neutralization of 25mL HCl (0.1 M) with NaOH (0.1 M) solution

Sl. No.	Volume of NaOH added (mL)	Total volume of NaOH added to the solution in flask (mL)	pH of solution	Change in pH
1.	0	0		–
2.	12.5	12.5		
3.	10.0	22.5		
4.	2.3	24.8		
5.	0.1	24.9		
6.	0.1	25.0		
7.	0.1	25.1		
8.	0.1	25.2		
9.	0.1	25.3		
10.	0.1	25.4		
11.	0.5	25.9		

Change in pH = (Higher value of pH) - (lower value of pH)

Plot a graph in graph paper (Page No. _____) :



Result :

(i) As 0.1 M NaOH solution is added in 25 mL 0.1 M HCl solution taken in flask, pH of solution _____ (increases/decreases)

(ii) Consequent pH values of total volume of _____ mL and _____ mL NaOH solution give higher value of change in pH. (Write consequent values of volumes from observation table)

Conclusion :

In initial stage of titration, pH value of solution is _____ as there is only HCl solution in flask. As NaOH solution is added from burette, pH of solution _____ (increases/decreases). After adding some volume of NaOH solution, by adding a very little amount (0.1 mL) of NaOH, _____ (very less increase/very high increase/very less decrease/very high decrease) in pH is observed. At this stage, total volume of NaOH indicate the end point of the titration.

- **Answer the following questions :**

1. What trend of pH change will you observe in the neutralization of strong acid with strong base ?

_____

_____

_____

2. Do you expect the same trend of pH change for neutralization of weak acid (acetic acid) with a strong base (sodium hydroxide) ?

_____

_____

3. In which pH range should the indicator show colour change if the hydrochloric acid is to be neutralized by sodium hydroxide? Give answer after looking at the graph of the experiment.

_____

_____

4. How does the study of pH change help in choosing the indicator for neutralization reaction. Explain.

_____

_____

_____

_____

Sign of subject teacher/Lab teacher

Date : _____

Experiment 11

Aim : To study pH of solutions of sodium chloride, ferric chloride and sodium carbonate.
[Reference: Experiment No. 5.5 of Laboratory Manual]

Apparatus Required :

Materials Required :

Procedure :

Observation :

In this experiment students can use either pH paper or universal indicator to measure the pH.

pH of NaCl, $FeCl_3$ and Na_2CO_3 solutions of different concentrations

Solution	pH of solutions (with pH paper or universal indicator)			
	Boiling Tube	Test Tube 1	Test Tube 2	Test Tube 3
NaCl				
$FeCl_3$				
Na_2CO_3				

Result :

(1) Aqueous solution of NaCl is _____ . (acidic/basic/neutral)

(2) Aqueous solution of $FeCl_3$ is _____ . (acidic/basic/neutral)

(3) Aqueous solution of Na_2CO_3 is _____ . (acidic/basic/neutral)

Conclusion :

(1) Aqueous solution of salt made from strong acid and strong base is _____ . (acidic/basic/neutral)

(2) Aqueous solution of salt made from strong acid and weak base is _____ . (acidic/basic/neutral)

(3) Aqueous solution of salt made from weak acid and strong base is _____ . (acidic/basic/neutral)

- **Answer the following questions :**

1. Why are $FeCl_3$ and Na_2CO_3 solutions not neutral ?

2. Why are the salts of strong acid and strong base not hydrolysed ? Explain.

3. How is the phenomenon of hydrolysis useful in salt analysis ?

4. What is the effect of dilution on pH of salt solution? Verify and explain your results.

Sign of subject teacher/Lab teacher

Date : _____

Experiment 12

Aim : To study the shift in equilibrium in the reaction of ferric ions and thiocyanate ions by increasing the concentration of any one of these ions.

[Reference: Experiment No. 4.1 of Laboratory Manual]

Apparatus Required :

Materials Required :

Procedure :

Observation :

Note down the observations due to increase in concentration of ferric ions and thiocyanate ions in following Table 1 and Table 2 respectively.

Table 1 : Equilibrium shift on increasing the concentration of ferric ions

Boiling Tube	Volume of ferric chloride solution taken in the system (mL)	Change in colour intensity as matched with reference solution in boiling tube "a"	Direction of shift in equilibrium
a	Reference solution for matching colour containing 2.5 mL blood red solution + 17.5 mL water (20 mL equilibrium mixture)		Equilibrium position
b	1.0		
c	2.0		
d	3.0		
e	4.0		

Table 2 : Equilibrium shift on increasing the concentration of thiocyanate ions

Boiling Tube	Volume of thiocyanate solution taken in the system (mL)	Change in colour intensity as matched with reference solution in boiling tube "a"	Direction of shift in equilibrium
a	Reference solution for matching colour containing 2.5 mL blood red solution + 17.5 mL water (20 mL equilibrium mixture)		Equilibrium position
b′	1.0		
c′	2.0		
d′	3.0		
e′	4.0		

Conclusion :

- **Answer the following questions :**

1. Explain why representing the ionic reaction between ferric and thiocyanate ions as given below $Fe^{3+}(aq) + SCN^{-}(aq) \rightleftharpoons [Fe(SCN)]^{2+}(aq)$ is more appropriate in the following form ?

   $[Fe(H_2O)_6]^{3+} + SCN^{-}(aq) \rightleftharpoons [Fe(H_2O)_5(SCN)]^{2+} + H_2O$

2. Does the constancy in colour intensity indicate the dynamic nature of equilibrium? Explain your answer with appropriate reasons.

3. What is equilibrium constant and how does it differ from the rate constant ?

4. It is always advisable to carry out the present experiment with dilute solutions. Why ?

5. What will be the effect of adding solid potassium chloride to the system at equilibrium? Verify your answer experimentally ?

6. Why boiling tubes of same size are used in the experiment ?

Sign of subject teacher/Lab teacher

Date : _____

Experiment 13

Aim : To study the shift in equilibrium in the reaction between $[Co(H_2O)_6]^{2+}$ and Cl^- ions, by changing the concentration of any one of these ions.

[Reference: Experiment No. 4.2 of Laboratory Manual]

Apparatus Required :

Materials Required :

Procedure :

Observation:

Note down the observations due to increase in amount of water and concentration of Cl^- ions in the following Table 1 and Table 2 respectively.

Table 1 : Shift in equilibrium on adding water

Sl. No.	Test tube	Volume of acetone added (mL)	Volume of $CoCl_2$ solution added (mL)	Volume of water added (mL)	Colour of mixture	Direction of shift in equilibrium
1.	A	1.0	3.0	0.0		
2.	B	0.8	3.0	0.2		
3.	C	0.6	3.0	0.4		
4.	D	0.4	3.0	0.6		
5.	E	0.2	3.0	0.8		

Table 2 : Shift in equilibrium on adding Cl^- ions

Sl. No.	Test tube	Volume of conc. HCl added (mL)	Volume of aquo complex solution added (mL)	Volume of water added (mL)	Colour of mixture	Direction of shift in equilibrium
1.	A′	0.5	1.5	2.0		
2.	B′	1.0	1.5	1.5		
3.	C′	1.5	1.5	1.0		
4.	D′	2.0	1.5	0.5		
5.	E′	2.5	1.5	0.0		

Conclusion :

- **Answer the following questions :**

1. What will be the effect of increasing the temperature of the reaction mixture at equilibrium ?

2. Can an aqueous solution of sodium chloride replace concentrated HCl ? Verify your answer experimentlly.

3. Why should the total volume of the solution in each test tube be kept same ?

Sign of subject teacher/Lab teacher

Date : _____

Experiment 14

Aim : To detect the positive and negative ions present in the given salt by qualitative analysis. [Reference: Experiment No. 7.1 of Laboratory Manual]

Apparatus Required :

_____

_____

[A] Dry tests of inorganic salt :

Sl. No.	Test	Observation	Inference
1.	Colour		
2.	Smell		
3.	Heated 0.5 g of salt in a dry test tube		
4.	Flame test		
5.	Borax bead test		
6.	Charcoal cavity test		
7.	Cobalt Nitrate test		

[B] Tests for detection of anion :

Sl. No.	Test	Observation	Inference
1.	0.1g salt + dilute H_2SO_4 on heating		

2.	0.1g salt + 1 mL concentrated H_2SO_4 on heating		
3.	Water extract or acidified sodium carbonate extract by dilute HCl + $BaCl_2$		
4.	Water extract of salt or sodium carbonate extract + acetic acid (till solution becomes acidic) + lead acetate		
5.	Water extract or sodium carbonate extract + concentrated HNO_3 + ammonium molybdate		

[C] Confirmatory tests of anion :

Sl. No.	Test	Observation	Inference

[D] Chemical equations of anion :

[E] Group tests for detection of cation :

Preparation of original solution : Original solution (O.S.) = Salt + _____ (Solvent)

Sl. No.	Test	Observation	Inference
1.	Salt + dilute NaOH solution, on heating		Group (0) is _____
2.	O.S. + dilute HCl		Group (I) is _____
3.	O.S. + dilute HCl + on passing H_2S gas		Group (II) is _____
4.	O.S. + heated after adding 2 - 3 drops of concentrated HNO_3 + 0.2 g NH_4Cl + add NH_4OH solution in excess (till it smells of ammonia)		Group (III) is _____
5.	O.S. + 0.2g NH_4Cl + NH_4OH (add till it smells of ammonia) + on passing H_2S gas		Group (IV) is _____
6.	O.S. + 0.2g NH_4Cl + NH_4OH (add till it smells of ammonia) + 0.5g $(NH_4)_2CO_3$		Group (V) is _____
7.	O.S. + 0.2g NH_4Cl + NH_4OH(add till it smells of ammonia) + Na_2HPO_4 solution		Group (VI) is _____

[F] Analysis of group _____ cations :

Sl. No.	Test	Observation	Inference

[G] Confirmatory tests of cation :

Sl. No.	Test	Observation	Inference

[H] Chemical equations of cation :

[I] Result table :

Cation of inorganic salt and its name	Anion of inorganic salt and its name	Molecular formula and name of inorganic salt

Sign of subject teacher/Lab teacher

Date : _____

Experiment 15

Aim : To detect the positive and negative ions present in the given salt by qualitative analysis. [Reference: Experiment No. 7.1 of Laboratory Manual]

Apparatus Required :

_____

_____

[A] Dry tests of inorganic salt :

Sl. No.	Test	Observation	Inference
1.	Colour		
2.	Smell		
3.	Heated 0.5 g of salt in a dry test tube		
4.	Flame test		
5.	Borax bead test		
6.	Charcoal cavity test		
7.	Cobalt Nitrate test		

[B] Tests for detection of anion :

Sl. No.	Test	Observation	Inference
1.	0.1g salt + dilute H_2SO_4 on heating		

2.	0.1g salt + 1 mL concentrated H_2SO_4 on heating		
3.	Water extract or acidified sodium carbonate extract by dilute HCl + $BaCl_2$		
4.	Water extract of salt or sodium carbonate extract + acetic acid (till solution becomes acidic) + lead acetate		
5.	Water extract or sodium carbonate extract + concentrated HNO_3 + ammonium molybdate		

[C] Confirmatory tests of anion :

Sl. No.	Test	Observation	Inference

[D] Chemical equations of anion :

[E] Group tests for detection of cation :

Preparation of original solution : Original solution (O.S.) = Salt + _____ (Solvent)

Sl. No.	Test	Observation	Inference
1.	Salt + dilute NaOH solution, on heating		Group (0) is _____
2.	O.S. + dilute HCl		Group (I) is _____
3.	O.S. + dilute HCl + on passing H_2S gas		Group (II) is _____
4.	O.S. + heated after adding 2 - 3 drops of concentrated HNO_3 + 0.2 g NH_4Cl + add NH_4OH solution in excess (till it smells of ammonia)		Group (III) is _____
5.	O.S. + 0.2g NH_4Cl + NH_4OH (add till it smells of ammonia) + on passing H_2S gas		Group (IV) is _____
6.	O.S. + 0.2g NH_4Cl + NH_4OH (add till it smells of ammonia) + 0.5g $(NH_4)_2CO_3$		Group (V) is _____
7.	O.S. + 0.2g NH_4Cl + NH_4OH(add till it smells of ammonia) + Na_2HPO_4 solution		Group (VI) is _____

[F] Analysis of group _____ cations :

Sl. No.	Test	Observation	Inference

[G] Confirmatory tests of cation :

Sl. No.	Test	Observation	Inference

[H] Chemical equations of cation :

[I] Result table :

Cation of inorganic salt and its name	Anion of inorganic salt and its name	Molecular formula and name of inorganic salt

Sign of subject teacher/Lab teacher

Date : _____

Experiment 16

Aim : To detect the positive and negative ions present in the given salt by qualitative analysis. [Reference: Experiment No. 7.1 of Laboratory Manual]

Apparatus Required :

_____

_____

[A] Dry tests of inorganic salt :

Sl. No.	Test	Observation	Inference
1.	Colour		
2.	Smell		
3.	Heated 0.5 g of salt in a dry test tube		
4.	Flame test		
5.	Borax bead test		
6.	Charcoal cavity test		
7.	Cobalt Nitrate test		

[B] Tests for detection of anion :

Sl. No.	Test	Observation	Inference
1.	0.1g salt + dilute H_2SO_4 on heating		

2.	0.1g salt + 1 mL concentrated H_2SO_4 on heating		
3.	Water extract or acidified sodium carbonate extract by dilute HCl + $BaCl_2$		
4.	Water extract of salt or sodium carbonate extract + acetic acid (till solution becomes acidic) + lead acetate		
5.	Water extract or sodium carbonate extract + concentrated HNO_3 + ammonium molybdate		

[C] Confirmatory tests of anion :

Sl. No.	Test	Observation	Inference

[D] Chemical equations of anion :

[E] Group tests for detection of cation :

Preparation of original solution : Original solution (O.S.) = Salt + _____ (Solvent)

Sl. No.	Test	Observation	Inference
1.	Salt + dilute NaOH solution, on heating		Group (0) is _____
2.	O.S. + dilute HCl		Group (I) is _____
3.	O.S. + dilute HCl + on passing H_2S gas		Group (II) is _____
4.	O.S. + heated after adding 2 - 3 drops of concentrated HNO_3 + 0.2 g NH_4Cl + add NH_4OH solution in excess (till it smells of ammonia)		Group (III) is _____
5.	O.S. + 0.2g NH_4Cl + NH_4OH (add till it smells of ammonia) + on passing H_2S gas		Group (IV) is _____
6.	O.S. + 0.2g NH_4Cl + NH_4OH (add till it smells of ammonia) + 0.5g $(NH_4)_2CO_3$		Group (V) is _____
7.	O.S. + 0.2g NH_4Cl + NH_4OH(add till it smells of ammonia) + Na_2HPO_4 solution		Group (VI) is _____

[F] Analysis of group _____ cations :

Sl. No.	Test	Observation	Inference

[G] Confirmatory tests of cation :

Sl. No.	Test	Observation	Inference

[H] Chemical equations of cation :

[I] Result table :

Cation of inorganic salt and its name	Anion of inorganic salt and its name	Molecular formula and name of inorganic salt

Sign of subject teacher/Lab teacher

Date : _____

Experiment 17

Aim : To detect the positive and negative ions present in the given salt by qualitative analysis. [Reference: Experiment No. 7.1 of Laboratory Manual]

Apparatus Required :

_____

_____

[A] Dry tests of inorganic salt :

Sl. No.	Test	Observation	Inference
1.	Colour		
2.	Smell		
3.	Heated 0.5 g of salt in a dry test tube		
4.	Flame test		
5.	Borax bead test		
6.	Charcoal cavity test		
7.	Cobalt Nitrate test		

[B] Tests for detection of anion :

Sl. No.	Test	Observation	Inference
1.	0.1g salt + dilute H_2SO_4 on heating		

2.	0.1g salt + 1 mL concentrated H_2SO_4 on heating		
3.	Water extract or acidified sodium carbonate extract by dilute HCl + $BaCl_2$		
4.	Water extract of salt or sodium carbonate extract + acetic acid (till solution becomes acidic) + lead acetate		
5.	Water extract or sodium carbonate extract + concentrated HNO_3 + ammonium molybdate		

[C] Confirmatory tests of anion :

Sl. No.	Test	Observation	Inference

[D] Chemical equations of anion :

[E] Group tests for detection of cation :

Preparation of original solution : Original solution (O.S.) = Salt + _____ (Solvent)

Sl. No.	Test	Observation	Inference
1.	Salt + dilute NaOH solution, on heating		Group (0) is _____
2.	O.S. + dilute HCl		Group (I) is _____
3.	O.S. + dilute HCl + on passing H_2S gas		Group (II) is _____
4.	O.S. + heated after adding 2 - 3 drops of concentrated HNO_3 + 0.2 g NH_4Cl + add NH_4OH solution in excess (till it smells of ammonia)		Group (III) is _____
5.	O.S. + 0.2g NH_4Cl + NH_4OH (add till it smells of ammonia) + on passing H_2S gas		Group (IV) is _____
6.	O.S. + 0.2g NH_4Cl + NH_4OH (add till it smells of ammonia) + 0.5g $(NH_4)_2CO_3$		Group (V) is _____
7.	O.S. + 0.2g NH_4Cl + NH_4OH(add till it smells of ammonia) + Na_2HPO_4 solution		Group (VI) is _____

[F] Analysis of group _____ cations :

Sl. No.	Test	Observation	Inference

[G] Confirmatory tests of cation :

Sl. No.	Test	Observation	Inference

[H] Chemical equations of cation :

[I] Result table :

Cation of inorganic salt and its name	Anion of inorganic salt and its name	Molecular formula and name of inorganic salt

Sign of subject teacher/Lab teacher

Date : _____

Experiment 18

Aim : To detect the positive and negative ions present in the given salt by qualitative analysis. [Reference: Experiment No. 7.1 of Laboratory Manual]

Apparatus Required :

_____

_____

[A] Dry tests of inorganic salt :

Sl. No.	Test	Observation	Inference
1.	Colour		
2.	Smell		
3.	Heated 0.5 g of salt in a dry test tube		
4.	Flame test		
5.	Borax bead test		
6.	Charcoal cavity test		
7.	Cobalt Nitrate test		

[B] Tests for detection of anion :

Sl. No.	Test	Observation	Inference
1.	0.1g salt + dilute H_2SO_4 on heating		

2.	0.1g salt + 1 mL concentrated H_2SO_4 on heating		
3.	Water extract or acidified sodium carbonate extract by dilute HCl + $BaCl_2$		
4.	Water extract of salt or sodium carbonate extract + acetic acid (till solution becomes acidic) + lead acetate		
5.	Water extract or sodium carbonate extract + concentrated HNO_3 + ammonium molybdate		

[C] Confirmatory tests of anion :

Sl. No.	Test	Observation	Inference

[D] Chemical equations of anion :

[E] Group tests for detection of cation :

Preparation of original solution : Original solution (O.S.) = Salt + _____ (Solvent)

Sl. No.	Test	Observation	Inference
1.	Salt + dilute NaOH solution, on heating		Group (0) is _____
2.	O.S. + dilute HCl		Group (I) is _____
3.	O.S. + dilute HCl + on passing H_2S gas		Group (II) is _____
4.	O.S. + heated after adding 2 - 3 drops of concentrated HNO_3 + 0.2 g NH_4Cl + add NH_4OH solution in excess (till it smells of ammonia)		Group (III) is _____
5.	O.S. + 0.2g NH_4Cl + NH_4OH (add till it smells of ammonia) + on passing H_2S gas		Group (IV) is _____
6.	O.S. + 0.2g NH_4Cl + NH_4OH (add till it smells of ammonia) + 0.5g $(NH_4)_2CO_3$		Group (V) is _____
7.	O.S. + 0.2g NH_4Cl + NH_4OH(add till it smells of ammonia) + Na_2HPO_4 solution		Group (VI) is _____

[F] Analysis of group _____ cations :

Sl. No.	Test	Observation	Inference

[G] Confirmatory tests of cation :

Sl. No.	Test	Observation	Inference

[H] Chemical equations of cation :

[I] Result table :

Cation of inorganic salt and its name	Anion of inorganic salt and its name	Molecular formula and name of inorganic salt

Sign of subject teacher/Lab teacher

Date : _____

Experiment 19

Aim : To detect the positive and negative ions present in the given salt by qualitative analysis. [Reference: Experiment No. 7.1 of Laboratory Manual]

Apparatus Required :

_____

_____

[A] Dry tests of inorganic salt :

Sl. No.	Test	Observation	Inference
1.	Colour		
2.	Smell		
3.	Heated 0.5 g of salt in a dry test tube		
4.	Flame test		
5.	Borax bead test		
6.	Charcoal cavity test		
7.	Cobalt Nitrate test		

[B] Tests for detection of anion :

Sl. No.	Test	Observation	Inference
1.	0.1g salt + dilute H_2SO_4 on heating		

2.	0.1g salt + 1 mL concentrated H_2SO_4 on heating		
3.	Water extract or acidified sodium carbonate extract by dilute HCl + $BaCl_2$		
4.	Water extract of salt or sodium carbonate extract + acetic acid (till solution becomes acidic) + lead acetate		
5.	Water extract or sodium carbonate extract + concentrated HNO_3 + ammonium molybdate		

[C] Confirmatory tests of anion :

Sl. No.	Test	Observation	Inference

[D] Chemical equations of anion :

[E] Group tests for detection of cation :

Preparation of original solution : Original solution (O.S.) = Salt + _____ (Solvent)

Sl. No.	Test	Observation	Inference
1.	Salt + dilute NaOH solution, on heating		Group (0) is _____
2.	O.S. + dilute HCl		Group (I) is _____
3.	O.S. + dilute HCl + on passing H_2S gas		Group (II) is _____
4.	O.S. + heated after adding 2 - 3 drops of concentrated HNO_3 + 0.2 g NH_4Cl + add NH_4OH solution in excess (till it smells of ammonia)		Group (III) is _____
5.	O.S. + 0.2g NH_4Cl + NH_4OH (add till it smells of ammonia) + on passing H_2S gas		Group (IV) is _____
6.	O.S. + 0.2g NH_4Cl + NH_4OH (add till it smells of ammonia) + 0.5g $(NH_4)_2CO_3$		Group (V) is _____
7.	O.S. + 0.2g NH_4Cl + NH_4OH(add till it smells of ammonia) + Na_2HPO_4 solution		Group (VI) is _____

[F] Analysis of group _____ cations :

Sl. No.	Test	Observation	Inference

[G] Confirmatory tests of cation :

Sl. No.	Test	Observation	Inference

[H] Chemical equations of cation :

[I] Result table :

Cation of inorganic salt and its name	Anion of inorganic salt and its name	Molecular formula and name of inorganic salt

Sign of subject teacher/Lab teacher

Date : _____

Experiment 20

Aim : To detect the positive and negative ions present in the given salt by qualitative analysis. [Reference: Experiment No. 7.1 of Laboratory Manual]

Apparatus Required :

_____

_____

[A] Dry tests of inorganic salt :

Sl. No.	Test	Observation	Inference
1.	Colour		
2.	Smell		
3.	Heated 0.5 g of salt in a dry test tube		
4.	Flame test		
5.	Borax bead test		
6.	Charcoal cavity test		
7.	Cobalt Nitrate test		

[B] Tests for detection of anion :

Sl. No.	Test	Observation	Inference
1.	0.1g salt + dilute H_2SO_4 on heating		

2.	0.1g salt + 1 mL concentrated H_2SO_4 on heating		
3.	Water extract or acidified sodium carbonate extract by dilute HCl + $BaCl_2$		
4.	Water extract of salt or sodium carbonate extract + acetic acid (till solution becomes acidic) + lead acetate		
5.	Water extract or sodium carbonate extract + concentrated HNO_3 + ammonium molybdate		

[C] Confirmatory tests of anion :

Sl. No.	Test	Observation	Inference

[D] Chemical equations of anion :

[E] Group tests for detection of cation :

Preparation of original solution : Original solution (O.S.) = Salt + _____ (Solvent)

Sl. No.	Test	Observation	Inference
1.	Salt + dilute NaOH solution, on heating		Group (0) is _____
2.	O.S. + dilute HCl		Group (I) is _____
3.	O.S. + dilute HCl + on passing H_2S gas		Group (II) is _____
4.	O.S. + heated after adding 2 - 3 drops of concentrated HNO_3 + 0.2 g NH_4Cl + add NH_4OH solution in excess (till it smells of ammonia)		Group (III) is _____
5.	O.S. + 0.2g NH_4Cl + NH_4OH (add till it smells of ammonia) + on passing H_2S gas		Group (IV) is _____
6.	O.S. + 0.2g NH_4Cl + NH_4OH (add till it smells of ammonia) + 0.5g $(NH_4)_2CO_3$		Group (V) is _____
7.	O.S. + 0.2g NH_4Cl + NH_4OH(add till it smells of ammonia) + Na_2HPO_4 solution		Group (VI) is _____

[F] Analysis of group _____ cations :

Sl. No.	Test	Observation	Inference

[G] Confirmatory tests of cation :

Sl. No.	Test	Observation	Inference

[H] Chemical equations of cation :

[I] Result table :

Cation of inorganic salt and its name	Anion of inorganic salt and its name	Molecular formula and name of inorganic salt

Sign of subject teacher/Lab teacher

Date : ____/____/____

Experiment 21

Aim : To detect the positive and negative ions present in the given salt by qualitative analysis. [Reference: Experiment No. 7.1 of Laboratory Manual]

Apparatus Required :

_____

_____

[A] Dry tests of inorganic salt :

Sl. No.	Test	Observation	Inference
1.	Colour		
2.	Smell		
3.	Heated 0.5 g of salt in a dry test tube		
4.	Flame test		
5.	Borax bead test		
6.	Charcoal cavity test		
7.	Cobalt Nitrate test		

[B] Tests for detection of anion :

Sl. No.	Test	Observation	Inference
1.	0.1g salt + dilute H_2SO_4 on heating		

2.	0.1g salt + 1 mL concentrated H_2SO_4 on heating		
3.	Water extract or acidified sodium carbonate extract by dilute HCl + $BaCl_2$		
4.	Water extract of salt or sodium carbonate extract + acetic acid (till solution becomes acidic) + lead acetate		
5.	Water extract or sodium carbonate extract + concentrated HNO_3 + ammonium molybdate		

[C] Confirmatory tests of anion :

Sl. No.	Test	Observation	Inference

[D] Chemical equations of anion :

[E] Group tests for detection of cation :

Preparation of original solution : Original solution (O.S.) = Salt + _____ (Solvent)

Sl. No.	Test	Observation	Inference
1.	Salt + dilute NaOH solution, on heating		Group (0) is _____
2.	O.S. + dilute HCl		Group (I) is _____
3.	O.S. + dilute HCl + on passing H_2S gas		Group (II) is _____
4.	O.S. + heated after adding 2 - 3 drops of concentrated HNO_3 + 0.2 g NH_4Cl + add NH_4OH solution in excess (till it smells of ammonia)		Group (III) is _____
5.	O.S. + 0.2g NH_4Cl + NH_4OH (add till it smells of ammonia) + on passing H_2S gas		Group (IV) is _____
6.	O.S. + 0.2g NH_4Cl + NH_4OH (add till it smells of ammonia) + 0.5g $(NH_4)_2CO_3$		Group (V) is _____
7.	O.S. + 0.2g NH_4Cl + NH_4OH(add till it smells of ammonia) + Na_2HPO_4 solution		Group (VI) is _____

[F] Analysis of group _____ cations :

Sl. No.	Test	Observation	Inference

[G] Confirmatory tests of cation :

Sl. No.	Test	Observation	Inference

[H] Chemical equations of cation :

[I] Result table :

Cation of inorganic salt and its name	Anion of inorganic salt and its name	Molecular formula and name of inorganic salt

- **Answer the following questions :**

1. What is the difference between a qualitative and a quantitative analysis?

2. Can we use glass rod instead of platinum wire for performing the flame test? Explain your answer.

3. Why is platinum metal preferred to other metals for the flame test?

4. Name the anions detected with the help of dilute H_2SO_4

5. Why is dilute H_2SO_4 preferred over dilute HCl while testing anions?

6. Name the anions detected by conc. H_2SO_4.

7. How is sodium carbonate extract prepared ?

8. What is lime water and what happens on passing carbon dioxide gas through it?

9. Carbon dioxide gas and sulphur dioxide gas both turn lime water milky. How will you distinguish between the two ?

10. How will you test the presence of carbonate ion ?

11. What is the composition of dark brown ring which is formed at the junction of two layers in the ring test for nitrates ?

12. Name the radical confirmed by sodium nitroprusside test.

13. What is chromyl chloride test ? How do you justify that CrO_2Cl_2 is acidic in nature ?

14. Why do bromides and iodides not give tests similar to chromyl chloride test ?

15. Describe the layer test for bromide and iodide ions.

16. Why is silver nitrate solution stored in dark coloured bottles ?

17. How do you test the presence of sulphide ion ?

18. Why does iodine give a blue colour with starch solution ?

19. What is Nessler's reagent ?

20. Why is original solution for cations not prepared in conc. HNO_3 or conc. H_2SO_4 ?

21. Why cannot conc. HCl be used as a group reagent in place of dil. HCl for the precipitation of first group cations ?

22. How can one prevent the precipitation of Group–IV radicals with the Group-II radicals ?

23. Why is it essential to boil off H_2S gas before precipitation of radicals of group–III ?

24. Why is heating with conc. nitric acid done before precipitation of group–III ?

25. Can we use ammonium sulphate instead of ammonium chloride in group–III ?

26. Why is NH_4OH added before $(NH_4)_2CO_3$ solution while precipitating group–V cations ?

27. Why do we sometimes get a white precipitate in group–VI even if the salt does not contain Mg^{2+} radical ?

28. What is aqua regia ?

29. Name a cation, which is not obtained from a metal.

30. How can you test the presence of ammonium ion ?

31. Why are the group–V radicals tested in the order Ba^{2+}, Sr^{2+} and Ca^{2+} ?

32. Why does conc. HNO_3 kept in a bottle turn yellow in colour?

33. Why should the solution be concentrated before proceeding to group–V ?

34. Why is the reagent bottle containing sodium hydroxide solution never stoppered ?

35. What do you understand by the term common ion effect ?

36. Why is zinc sulphide not precipitated in group–II ?

Sign of subject teacher/Lab teacher

Date : _____

Experiment 22

Aim : To purify compound by crystallization.

[Reference: Experiment No. 3.1 of Laboratory Manual]

Apparatus Required :

Materials Required :

Name of compound taken as sample :

Procedure :

Observation :

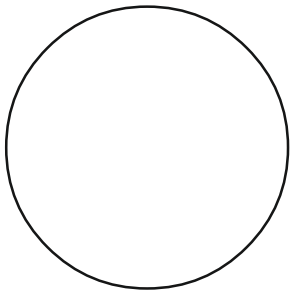
(1) Do you observe difference in external appearance of the compound before and after crystallization ? _____ (yes / No)

(2) Colour of crystals : _____ .

(3) Size and shape of crystals : _____ .

(4) Do the obtained crystals exhibit luster? _____ . (yes / No)

(5) Draw the figure of obtained crystals in given circle as per your observations.



Obtained Crystals

Result :

- **Answer the following questions :**

1. Which of the following formula is correct representation of potash alum (phitkari) ? Explain.

   (a) $K^{+}(H_2O)_6Al^{3+}(H_2O)_6(SO_4^{2-})_2$

   (b) $K_2SO_4 \cdot Al_2(SO_4)_3 \cdot 24H_2O$

2. What are isomorphous compounds ?

3. What is meant by the term, 'water of crystallisation' ?

4. Describe the effect of strong heating on each type of crystal prepared by you.

5. What do you understand by the term 'mother liquor' ?

6. Which thermodynamic function favours the process of crystallization ?

7. Explain the term – saturated solution.

8. Why is the preparation of saturated solution essential for making crystals ?

9. Name the processes involved in crystallization.

10. What is Kipp's waste? How can we obtain crystals of ferrous sulphate from Kipp's waste ?

Sign of subject teacher/Lab teacher

Date : ____/____/____

Experiment 23

Aim : Determination of melting point of a solid organic compound.

[Reference: Experiment No. 3.2 of Laboratory Manual]

Apparatus Required :

Materials Required :

Name of compound taken as sample :

Procedure :

(**Note:** Different apparatuses are used for determination of melting point. Out of them, use of beaker is easier method to determine melting point. Arrangement of apparatus and material for experiment is made according to beaker containing method given in Figure 3.1 (b) of Laboratory Manual)

Draw figure of arrangement of apparatus and material

Observation:

Mercury level in thermometer when compound of capillary tube melts = _____ ^{o}C

Result:

Melting point of given solid organic compound is _____ ^{o}C.

- **Answer the following questions :**

1. Why do pure solids possess sharp melting point ?

2. What is the effect of impurities on the melting point of a solid ?

3. Why is the melting point of benzamide more than acetamide ?

4. Can any other liquid be used in place of liquid paraffin to determine the melting point ?

5. Can we heat the capillary directly for the determination of melting point ?

6. Why is liquid paraffin filled in the Thiele's tube/Kjeldahl's flask ?

7. Why is Thiele's tube heated at the side arm ?

Sign of subject teacher/Lab teacher

Date : _____

Experiment 24

Aim : Determination of boiling point of a liquid organic compound.

[Reference: Experiment No. 3.3 of Laboratory Manual]

Apparatus Required :

Materials Required :

Name of compound taken as sample :

Procedure :

(**Note:** Different apparatuses are used for determination of boiling point. Out of them, use of beaker is easier method to determine boiling point. Arrangement of apparatus and material for experiment is made according to beaker containing method instead of ignition tube given in Figure 3.2 of Laboratory Manual)

Draw figure of arrangement of apparatus and material

Observation:

Mercury level in thermometer when bubbles start vigorusly at the lower end of the capillary dipped in the liquid organic compound continuously = _____ ^{o}C.

Result:

Boiling point of given liquid organic compound is _____ °C.

- **Answer the following questions :**

1. Suggest a suitable liquid, which can be filled in the Thiele's tube for the determination of the boiling point of carbon tetrachloride.

_____

_____

_____

2. In place of liquid paraffine, can any other liquid be used for the purpose of determination of boiling point ?

_____

_____

_____

3. Suppose boiling point of a liquid in Delhi is 100°C. At hill station, will it be the same or different? Give reasons.

_____

_____

_____

4. Why is food cooked more quickly in a pressure cooker ?

_____

_____

5. How would the boiling point of water vary with the addition of equimolar quantities of urea, potassium chloride and potassium sulphate ?

_____

_____

6. Why do different isomers of alcohol represented by the formula $C_4H_{10}O$ differ in their boiling points ?

_____

_____

_____

Sign of subject teacher/Lab teacher

Experiment No. :.....

Date :.....

Scale

X-axis = ...

Y-axis = ...

Experiment No. :.....

Date :.....

Scale

X-axis = ...

Y-axis = ...

LOGARITHMS

Table 1

N	0	1	2	3	4	5	6	7	8	9	1	2	3	4	5	6	7	8	9
10	0000	0043	0086	0128	0170	0212	0253	0294	0334	0374	5	9	13	17	21	26	30	34	38
											4	8	12	16	20	24	28	32	36
11	0414	0453	0492	0531	0569	0607	0645	0682	0719	0755	4	8	12	16	20	23	27	31	35
											4	7	11	15	18	22	26	29	33
12	0792	0828	0864	0899	0934	0969	1004	1038	1072	1106	3	7	11	14	18	21	25	28	32
											3	7	10	14	17	20	24	27	31
13	1139	1173	1206	1239	1271	1303	1335	1367	1399	1430	3	6	10	13	16	19	23	26	29
											3	7	10	13	16	19	22	25	29
14	1461	1492	1523	1553	1584	1614	1614	1673	1703	1732	3	6	9	12	15	19	22	25	28
											3	6	9	12	14	17	2	023	26
15	1791	1790	1818	1847	1875	1903	1931	1959	1987	2014	3	6	9	11	14	17	20	23	26
											3	6	8	11	14	17	19	22	25
16	2041	2068	2098	2122	2148	2175	2201	2227	2253	2279	3	6	8	11	14	16	19	22	24
											3	5	8	10	13	16	18	21	23
17	2304	2330	2365	2380	2405	2430	2455	2480	2504	2529	3	5	8	10	13	15	18	20	23
											3	5	8	10	12	15	17	20	22
18	2553	2577	2601	2625	2648	2672	2695	2718	2742	2765	2	5	7	9	12	14	17	19	21
											2	4	7	9	11	14	16	18	21
19	2788	2810	2833	2856	2878	2900	2923	2945	2967	2989	2	4	7	9	11	13	16	18	20
											2	4	6	8	11	13	15	17	19
20	3010	3032	3054	3075	3096	3118	3139	3160	3181	3201	2	4	6	8	11	13	15	17	19
21	3222	3243	3263	3284	3304	3324	3345	3365	3385	3404	2	4	6	8	10	12	14	16	18
22	3424	3444	3464	3483	3502	3522	3540	3560	3579	3598	2	4	6	8	10	12	14	15	17
23	3617	3636	3655	3675	3692	3711	3729	3747	3766	3784	2	4	6	7	9	11	13	15	17
24	3802	3820	3838	3856	3874	3892	3909	3927	3945	3962	2	4	5	7	9	11	12	14	16
25	3979	3997	4014	4031	4048	4065	4082	4099	4116	4133	2	3	5	7	9	10	12	14	15
26	4150	4166	4182	4200	4216	4232	4279	4265	4281	4298	2	3	5	7	8	10	11	13	15
27	4314	4330	4346	4362	4378	4393	4409	4425	4440	4456	2	3	5	6	8	9	11	13	14
28	4472	4487	4502	4518	4533	1548	4564	4579	4594	4609	2	3	5	6	8	9	11	12	14
29	4624	4639	4654	4669	4683	4698	4713	4728	4742	4757	1	3	4	6	7	9	10	12	13
30	4771	4786	4800	4814	4829	4843	4857	4871	4886	4900	1	3	4	6	7	9	10	11	13
31	4914	4928	4942	4955	4969	4983	4997	5011	5024	5038	1	3	4	6	7	8	10	11	12
32	5051	5065	5079	5092	5105	5119	5132	5142	5159	5172	1	3	4	5	7	8	9	11	12
33	5185	5198	5211	5224	5238	5250	5263	5276	5289	5302	1	3	4	5	6	8	9	10	12
34	5315	5328	5340	5353	5366	5378	5391	5403	5416	5428	1	3	4	5	6	8	9	10	11
35	5441	5453	5465	5478	5490	5502	5514	5527	5539	5551	1	2	4	5	6	7	9	10	11
36	5563	5575	5587	5599	5611	5623	5635	5647	5658	5670	1	2	4	5	6	7	8	10	11
37	5682	5694	5705	5717	5729	5740	5752	5763	5775	5786	1	2	3	5	6	7	8	9	10
38	5798	5809	5821	532	5843	5855	5866	5877	5888	5899	1	2	3	5	6	7	8	9	10
39	5911	5922	5933	5944	5922	5966	5977	5988	5999	6010	1	2	3	4	5	7	8	9	10
40	6021	6031	6042	6053	6064	6075	6085	6096	6107	6117	1	2	3	4	5	6	8	9	10
41	6128	6138	6149	6160	6170	6180	6191	6201	6212	6222	1	2	3	4	5	6	7	8	9
42	6232	6243	6253	6263	6274	6284	6294	6304	6314	6325	1	2	3	4	5	6	7	8	9
43	6335	6345	6355	6365	6375	6385	6395	6405	6415	6425	1	2	3	4	5	6	7	8	9
44	6435	6444	6454	6464	6474	6484	6493	6503	6513	6522	1	2	3	4	5	6	7	8	9
45	6235	6542	6551	6561	6571	6580	6590	6599	6609	6618	1	2	3	4	5	6	7	8	9
46	6628	6637	6646	6656	6665	6675	6684	6693	6702	6712	1	2	3	4	5	6	7	7	8
47	6721	6730	6739	6749	6758	6767	6776	6785	6794	6803	1	2	3	4	5	5	6	7	8
48	6812	6821	6830	6839	6848	6857	6866	6875	6884	6893	1	2	3	4	4	5	6	7	8
49	6902	6911	6920	6928	6937	6946	6955	6964	6972	6981	1	2	3	4	4	5	6	7	8

Table 1

N	0	1	2	3	4	5	6	7	8	9	1	2	3	4	5	6	7	8	9
50	6990	6998	7007	7016	7024	7033	7042	7050	7059	7067	1	2	3	3	4	5	6	7	8
51	7076	7084	7093	7101	7110	7118	7126	7135	7143	7152	1	2	3	3	4	5	6	7	8
52	7160	7168	7177	7185	7193	7202	7210	7218	7226	7235	1	2	2	3	4	5	6	7	7
53	7243	7251	7259	7267	7275	7284	7292	7300	7308	7316	1	2	2	3	4	5	6	6	7
54	7324	7332	7340	7348	7356	7364	7372	7380	7388	7396	1	2	2	3	4	5	6	6	7
55	7404	7412	7419	7127	7435	7443	7451	7459	7466	7474	1	2	2	3	4	5	5	6	7
56	7482	7490	7497	7505	7513	7530	7528	7536	7543	7551	1	2	2	3	4	5	5	6	7
57	7559	7566	7574	7582	7589	7597	7604	7612	7619	7627	1	2	2	3	4	5	5	6	7
58	7634	7642	7649	7657	7664	7672	7679	7686	7694	7701	1	1	2	3	4	4	5	6	7
59	7709	7716	7723	7731	7738	7745	7752	7760	7767	7774	1	1	2	3	4	4	5	6	7
60	7782	7789	7796	7803	7810	7818	7825	7832	7839	7846	1	1	2	3	4	4	5	6	6
61	7853	7860	7768	7875	7882	7889	7896	7903	7910	7917	1	1	2	3	4	4	5	6	6
62	7924	7931	7938	7945	7952	7959	7966	7973	7980	7987	1	1	2	3	3	4	5	6	6
63	7993	8000	8007	8014	8021	8028	8035	8041	8048	8055	1	1	2	3	3	4	5	6	6
64	8062	8069	8075	8082	8089	8096	8102	8109	8116	8122	1	1	2	3	3	4	5	5	6
65	8129	8136	8142	8149	8156	8162	8169	8176	8182	8189	1	1	2	3	3	4	5	5	6
66	8195	8202	8209	8215	8222	8228	8235	8241	8248	8254	1	1	2	3	3	4	5	5	6
67	8261	8267	8274	8280	8287	8293	8299	8306	8312	8319	1	1	2	3	3	4	5	5	6
68	8325	8331	8338	8344	8351	8357	8367	8370	8376	8382	1	1	2	3	3	4	4	5	6
69	8388	8395	8401	8407	8414	8420	8426	8432	8439	8445	1	1	2	2	3	4	4	5	6
70	8451	8457	8463	8470	8476	8482	8488	8494	8500	8506	1	1	2	2	3	4	4	5	6
71	8513	8519	8525	8531	8537	8543	8549	8555	8561	8567	1	1	2	2	3	4	4	5	5
72	8573	8579	8585	8591	8597	8603	8609	8615	8621	8627	1	1	2	2	3	4	4	5	5
73	8633	8639	8645	8651	8657	8663	8669	8675	8681	8686	1	1	2	2	3	4	4	5	5
74	8692	8698	8704	8710	8716	8722	8727	8733	8739	8745	1	1	2	2	3	4	4	5	5
75	8751	8756	8762	8768	8774	8779	8785	8791	8797	8802	1	1	2	2	3	3	4	5	5
76	8808	8814	8820	8825	8831	8837	8842	8848	8854	8859	1	1	2	2	3	2	4	5	6
77	8865	8871	8876	8882	8887	8893	8899	8904	8910	8915	1	1	2	2	3	3	4	4	5
78	8921	8927	8932	8938	8943	8949	8954	8960	8965	8971	1	1	2	2	3	3	4	4	5
79	8976	8982	8987	8993	8998	9004	9009	9015	9020	9025	1	1	2	2	3	3	4	4	5
80	9031	9036	9042	9047	9053	9058	9063	9069	9074	9079	1	1	2	2	3	3	4	4	5
81	9085	9090	9096	9101	9106	9112	9117	9122	9128	9133	1	1	2	2	3	3	4	4	5
82	9138	9143	9149	9154	9159	9165	9170	9175	9180	9186	1	1	2	2	3	3	4	4	5
83	9191	9196	9201	9206	9212	9217	9222	9227	9232	9238	1	1	2	2	3	3	4	4	5
84	9243	9248	9253	9258	9263	9267	9274	9279	9284	9289	1	1	2	2	3	3	4	4	5
85	9294	9299	9304	9309	9315	9320	9325	9330	9335	9340	1	1	2	2	3	3	4	4	5
86	9345	9350	9355	9360	9365	9370	9375	9380	9385	9390	1	1	2	2	3	3	4	4	5
87	9395	9400	9405	9410	9415	9420	9425	9430	9435	9440	0	1	1	2	2	3	3	4	4
88	9445	9450	9455	9460	9465	9469	9474	9479	9484	9489	0	1	1	2	2	3	3	4	4
89	9494	9499	9504	9509	9513	9518	9523	9528	9533	9538	0	1	1	2	2	3	3	4	4
90	9542	9547	9552	9557	9562	9566	9571	9576	9581	9586	0	1	1	2	2	3	3	4	4
91	9590	9595	9600	9605	9609	9614	9619	9624	9628	9633	0	1	1	2	2	3	3	4	4
92	9638	9643	9647	9652	9657	9661	9666	9671	9675	9680	0	1	1	2	2	3	3	4	4
93	9685	9689	9694	9699	9703	9708	9713	9717	9722	9727	0	1	1	2	2	3	3	4	4
94	9731	9736	9741	9745	9750	9754	9759	9763	9768	9773	0	1	1	2	2	3	3	4	4
95	9777	9782	9786	9791	9795	9800	9805	9809	9814	9818	0	1	1	2	2	3	3	4	4
96	9823	9827	9832	9836	9841	9845	9850	9854	9859	9863	0	1	1	2	2	3	3	4	4
97	9868	9872	9877	9881	9886	9890	9894	9899	9903	9908	0	1	1	2	2	3	3	4	4
98	9912	9917	9921	9926	9930	9934	9939	9943	9948	9952	0	1	1	2	2	3	3	4	4
99	9956	9961	9965	9969	9974	9978	9983	9987	9997	9996	0	1	1	2	2	3	3	3	4

ANTILOGARITHMS

Table 2

N	0	1	2	3	4	5	6	7	8	9	1	2	3	4	5	6	7	8	9
.00	1000	1002	1005	1007	1009	1012	1014	1016	1019	1021	0	0	1	1	1	1	2	2	2
.01	1023	1026	1028	1030	1033	1035	1038	1040	1042	1045	0	0	1	1	1	1	2	2	2
.02	1047	1050	1052	1054	1057	1059	1062	1064	1067	1069	0	0	1	1	1	1	2	2	2
.03	1072	1074	1076	1079	1081	1084	1086	1089	1091	1094	0	0	1	1	1	1	2	2	2
.04	1096	1099	1102	1104	1107	1109	1112	1114	1117	1119	0	1	1	1	1	2	2	2	2
.05	1122	1125	1127	1130	1132	1135	1138	1140	1143	1146	0	1	1	1	1	2	2	2	2
.06	1148	1151	1153	1156	1159	1161	1164	1167	1169	1172	0	1	1	1	1	2	2	2	2
.07	1175	1178	1180	1183	1186	1189	1191	1194	1197	1199	0	1	1	1	1	2	2	2	2
.08	1202	1205	1208	1211	1213	1216	1219	1222	1225	1227	0	1	1	1	1	2	2	2	3
.09	1230	1233	1236	1239	1242	1245	1247	1250	1253	1256	0	1	1	1	1	2	2	2	3
.10	1259	1262	1265	1268	1271	1274	1276	1279	1282	1285	0	1	1	1	1	2	2	2	3
.11	1288	1291	1294	1297	1300	1303	1306	1309	1312	1315	0	1	1	1	2	2	2	2	3
.12	1318	1321	1324	1327	1330	1334	1337	1340	1343	1346	0	1	1	1	2	2	2	2	3
.13	1349	1352	1355	1358	1361	1365	1368	1371	1374	1377	0	1	1	1	2	2	2	3	3
.14	1380	1384	1387	1390	1393	1396	1400	1403	1406	1409	0	1	1	1	2	2	2	3	3
.15	1413	1416	1419	1422	1426	1429	1432	1435	1439	1442	0	1	1	1	2	2	2	3	3
.16	1445	1449	1452	1455	1459	1462	1466	1469	1472	1479	0	1	1	1	2	2	2	3	3
.17	1479	1483	1486	1489	1493	1496	1500	1503	1507	1510	0	1	1	1	2	2	2	3	3
.18	1514	1517	1521	1524	1528	1531	1535	1538	1542	1545	0	1	1	1	2	2	2	3	3
.19	1549	1552	1556	1560	1563	1567	1570	1574	1578	1581	0	1	1	1	2	2	3	3	3
.20	1585	1589	1592	1596	1600	1603	1607	1611	1614	1618	0	1	1	1	2	2	3	3	3
.21	1622	1626	1629	1633	1637	1641	1644	1648	1652	1656	0	1	1	2	2	2	3	3	3
.22	1660	1663	1667	1671	1675	1679	1683	1687	1690	1694	0	1	1	2	2	2	3	3	3
.23	1698	1702	1706	1710	1714	1718	1722	1726	1730	1734	0	1	1	2	2	2	3	3	4
.24	1738	1742	1746	1750	1754	1758	1762	1766	1770	1774	0	1	1	2	2	2	3	3	4
.25	1778	1782	1786	1791	1795	1799	1803	1807	1811	1816	0	1	1	2	2	2	3	3	4
.26	1820	1824	1828	1832	1837	1841	1845	1849	1854	1858	0	1	1	2	2	3	3	3	4
.27	1862	1866	1871	1875	1879	1884	1888	1892	1897	1901	0	1	1	2	2	3	3	3	4
.28	1905	1910	1914	1919	1923	1928	1932	1936	1941	1945	0	1	1	2	2	3	3	4	4
.29	1950	1954	1959	1963	1968	1972	1977	1982	1986	1991	0	1	1	2	2	3	3	4	4
.30	1995	2000	2004	2009	2014	2018	2023	2028	2032	2037	0	1	1	2	2	3	3	4	4
.31	2042	2046	2051	2056	2061	2065	2070	2075	2080	2084	0	1	1	2	2	3	3	4	4
.32	2089	2094	2099	2104	2109	2113	2118	2123	2128	2133	0	1	1	2	2	3	3	4	4
.33	2138	2143	2148	2153	2158	2163	2168	2173	2178	2183	0	1	1	2	2	3	3	4	4
.34	2188	2193	2198	2203	2208	2213	2218	2223	2228	2234	1	1	2	2	3	3	4	4	5
.35	2239	2244	2249	2254	2259	2265	2270	275	2280	2286	1	1	2	2	3	3	4	4	5
.36	2291	2296	2301	2307	2312	2317	2323	2328	2333	2339	1	1	2	2	3	3	4	4	5
.37	2344	2350	2355	2360	2366	2371	2377	2382	2388	2393	1	1	2	2	3	3	4	4	5
.38	2399	2404	2410	2415	2421	2427	2432	2438	2443	2449	1	1	2	2	3	3	4	4	5
.39	2455	2460	2466	2472	2477	2483	2489	2495	2500	2506	1	1	2	2	3	3	4	5	5
.40	2512	2518	2523	2529	2535	2541	2547	2553	2559	2564	1	1	2	2	3	4	4	5	5
.41	2570	2576	2582	2588	2594	2600	2606	2612	2618	2624	1	1	2	2	3	4	4	5	5
.42	2630	2636	2642	2649	2655	2661	2667	2673	2679	2685	1	1	2	2	3	4	4	5	6
.43	2692	2698	2704	2710	2716	2723	2729	2735	2742	2748	1	1	2	3	3	4	4	5	6
.44	2754	2761	2767	2773	2780	2786	2793	2799	2805	2812	1	1	2	3	3	4	4	5	6
.45	2818	2825	2831	2838	2844	2851	2858	2864	2871	2877	1	1	2	3	3	4	5	5	6
.46	2884	2891	2897	2904	2911	2917	2924	2931	2938	2944	1	1	2	3	3	4	5	5	6
.47	2951	2958	2965	2972	2979	2985	2992	2999	3006	3013	1	1	2	3	3	4	5	5	6
.48	3020	3027	3034	3041	3048	3055	3062	3069	3076	3083	1	1	2	3	3	4	5	6	6
.49	3090	3097	3105	3112	3119	3126	3133	3141	3148	3155	1	1	2	3	3	4	5	6	6

Table 2

N	0	1	2	3	4	5	6	7	8	9	1	2	3	4	5	6	7	8	9
.50	3162	3170	3177	3184	3192	3199	3206	3214	3221	3228	1	1	2	3	4	4	5	6	7
.51	3236	3243	3251	3258	3266	3273	3281	3289	3296	3304	1	2	2	3	4	5	5	6	7
.52	3311	3319	3327	3334	3342	3350	3357	3365	3373	3381	1	2	2	3	4	5	5	6	7
.53	3388	3396	3404	3412	3420	3428	3436	3443	3451	3459	1	2	2	3	4	5	6	6	7
.54	3467	3475	3483	3491	3499	3508	3513	3524	3532	3540	1	2	2	3	4	5	6	6	7
.55	3548	3556	3565	3573	3581	3589	3597	3606	3614	3622	1	2	2	3	4	5	6	7	7
.56	3631	3639	3648	3656	3664	3673	3681	3690	3698	3707	1	2	3	3	4	5	6	7	8
.57	3715	3724	3733	3741	3750	3758	3767	3776	3784	3793	1	2	3	3	4	5	6	7	8
.58	3802	3811	3819	3828	3837	3846	3855	3864	3873	3882	1	2	3	4	4	5	6	7	8
.59	3890	3899	3908	3917	3926	3936	3945	3954	3963	3972	1	2	3	4	5	5	6	7	8
.60	3981	3990	3999	4009	4018	4027	4036	4046	4055	4064	1	2	3	4	5	6	6	7	8
.61	4074	4083	4093	4102	4111	4121	4130	4140	4150	4159	1	2	3	4	5	6	7	8	9
.62	4169	4178	4188	4198	4207	4217	4227	4236	4246	4256	1	2	3	4	5	6	7	8	9
.63	4266	4276	4285	4295	4305	4315	4325	4335	4345	4355	1	2	3	4	5	6	7	8	9
.64	4365	4375	4385	4395	4406	4416	4426	4436	4446	4457	1	2	3	4	5	6	7	8	9
.65	4467	4477	4487	4498	4508	4519	4529	4539	4550	4560	1	2	3	4	5	6	7	8	9
.66	4571	4581	4592	4603	4613	4624	4634	4645	4656	4667	1	2	3	4	5	6	7	9	10
.67	4677	4688	4699	4710	4721	4732	4742	4753	4764	4775	1	2	3	4	5	7	8	9	10
.68	4786	4797	4808	4819	4831	4842	4853	4864	4875	4887	1	2	3	4	5	7	8	9	10
.69	4898	4909	4920	4932	4943	4955	4966	4977	4989	5000	1	2	3	4	5	7	8	9	10
.70	5012	5023	5035	5047	5058	5070	5082	5093	5105	5117	1	2	4	5	6	7	8	9	11
.71	5129	5140	5152	5164	5176	5188	5200	5212	5224	5236	1	2	4	5	6	7	8	10	11
.72	5248	5260	5272	5284	5297	5309	5321	5333	5346	5358	1	2	4	5	6	7	9	10	11
.73	5370	5383	5395	5408	5420	5433	5445	5458	5470	5483	1	3	4	5	6	8	9	10	11
.74	5495	5508	5521	5534	5546	5559	5572	5585	5598	5610	1	3	4	5	6	8	9	10	12
.75	5623	5636	5649	5662	5675	5689	5702	5715	5728	5741	1	3	4	5	7	8	9	10	12
.76	5754	5768	5781	5794	5808	5821	5834	5848	5861	5875	1	3	4	5	7	8	9	11	12
.77	5888	5902	5916	5929	5943	5957	5970	5984	5998	6012	1	3	4	5	7	8	10	11	12
.78	6026	6039	6053	6067	6082	6095	6109	6124	6138	6152	1	3	4	6	7	8	10	11	13
.79	6166	6180	6194	6209	6223	6237	6252	6266	6281	6295	1	3	4	6	7	9	10	11	13
.80	6310	6324	6339	6353	6368	6383	6397	6412	6427	6442	1	3	4	6	7	9	10	12	13
.81	6457	6471	6486	6501	6516	6531	6546	6561	6577	6592	2	3	5	6	8	9	11	12	14
.82	6607	6622	6637	6653	6668	6683	6699	6714	6730	6745	2	3	5	6	8	9	11	12	14
.83	6761	6776	6792	6808	6823	6839	6855	6871	6887	6902	2	3	5	6	8	9	11	13	14
.84	6918	6934	6950	6966	6982	6998	7015	7031	7047	7063	2	3	5	6	8	10	11	13	15
.85	7079	7096	7112	7129	7145	761	7178	7194	7211	7228	2	3	5	7	8	10	12	13	15
.86	7244	7261	7278	7295	7311	7328	7345	7362	7379	7396	2	3	5	7	8	10	12	13	15
.87	7413	7430	7447	7464	7482	7499	7516	7534	7551	7568	2	3	5	7	9	10	12	14	16
.88	7586	7603	7621	7638	7656	7674	7691	7709	7727	7745	2	4	5	7	9	11	12	14	16
.89	7765	7780	7798	7816	7834	7852	7870	7889	7907	7925	2	4	5	7	9	11	13	14	16
.90	7943	7962	7980	7998	8017	8035	8054	8072	8091	8110	2	4	6	7	9	11	13	15	17
.91	8128	8147	8166	8185	8204	8222	8241	8260	8279	8299	2	4	6	8	9	11	13	15	17
.92	8318	8337	8356	8375	8395	8414	8433	8453	8472	8492	2	4	6	8	10	12	14	15	17
.93	8511	8531	8551	8570	8590	8610	8630	8650	8670	8690	2	4	6	8	10	12	14	16	18
.94	8710	8730	8750	8770	8790	8810	8831	8851	8872	8892	2	4	6	8	10	12	14	16	18
.95	8913	8933	8954	8974	8995	9016	9036	9057	9078	9099	2	4	6	8	10	12	15	17	19
.96	9120	9141	9162	9186	9204	9226	9247	9268	9290	9311	2	4	6	8	11	13	15	17	19
.97	9333	9354	9376	9397	9419	9441	9462	9484	9506	9525	2	4	7	9	11	13	15	17	20
.98	9550	9575	9594	9616	9638	9661	9683	9705	9727	9750	2	4	7	9	11	13	16	18	20
.99	9772	9795	9817	9840	9863	9886	9908	9931	9954	9977	2	5	7	9	11	14	16	18	20